

Notes from the Oesper Collections

The Chemical Uses of Clay Tobacco Pipes

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Among the thousands of items in the Jensen-Thomas Apparatus Collection is a small selection of clay tobacco pipes (figure 1). This particular style of pipe, in both its short- and long-stem variations, dates back at least to the late 16th century and is most often associated with the countries of Northern Europe and Great Britain. Less well-known is the fact that this simple common everyday object was often employed by chemists, from the late 18th century through the first half of the 20th century, as an inexpensive substitute for more elaborate and expensive chemical apparatus – whence their presence in our collections.

One use of such pipes in the chemical laboratory was as a substitute for more expensive crucibles when carrying out high-temperature reactions. An example dating from 1947 is shown in figure 2, where the pipe is used as a vessel for the high-temperature electrolysis of molten salts, such as potassium hydroxide (2):

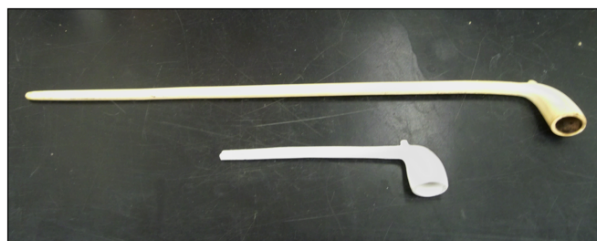
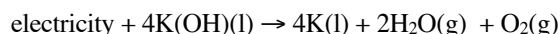


Figure 1. Typical examples of clay tobacco pipes found in the Jensen-Thomas Apparatus Collection.

A second example is shown in figure 3, which dates from 1928 and shows the destructive distillation of coal to produce coal gas – a complex mixture of methane, carbon monoxide and other hydrocarbons which was used in gas lamps and gas stoves throughout the 19th and early 20th centuries. In this experiment the bowl of the pipe has been sealed using a clay plug and the pipe functions as a substitute for a special form of crucible known as a Rose crucible.

Yet a third example, dating from 1925, is shown in figure 4, this time involving a long-stem rather than a

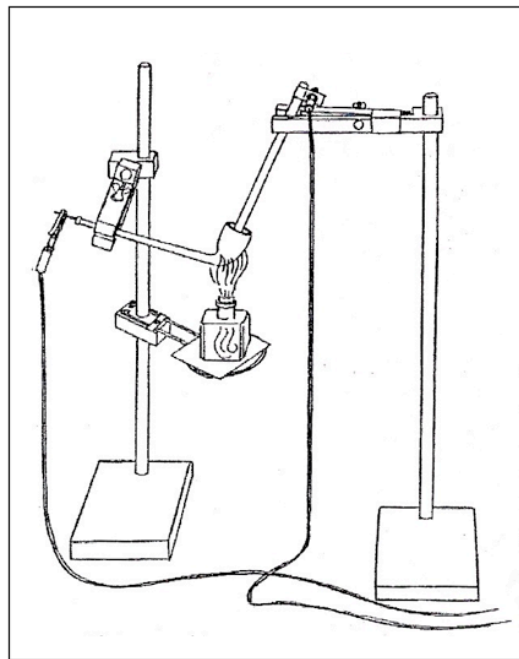


Figure 2. Electrolytic preparation of potassium using a clay pipe as the reaction vessel.

short-stem version of the pipe (4). Here the pipe-stem is acting as a substitute for a ceramic combustion tube and is being used to effect the thermal decomposition of nitric acid:

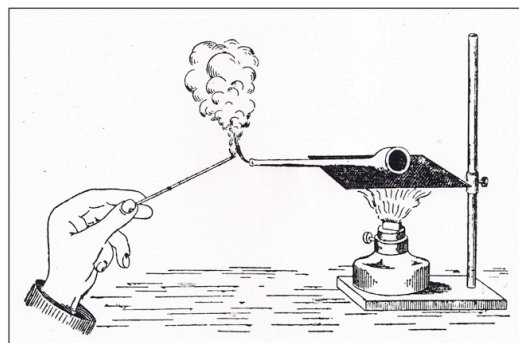
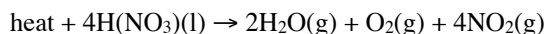


Figure 3. A “miniature gas works.”

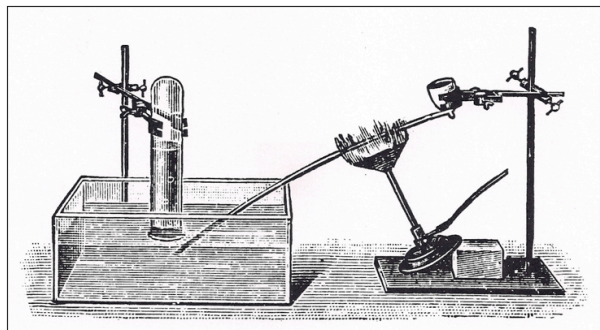
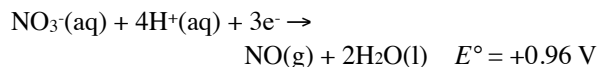


Figure 4. The thermal decomposition of nitric acid.

The bowl of the pipe has been retained as a convenient funnel for introducing a small quantity of nitric acid into the pipe stem.

Clay tobacco pipes were also used as substitutes for the porous membranes in voltaic cells, as shown in figure 5, where the broken-off bowl of the pipe serves as the separator for a small Grove cell, first introduced by the British chemist, William Grove, in 1839 (5). Here the bowl, whose stem hole has been blocked with a clay plug, acts as the cathode compartment and contains both an inert Pt electrode and concentrated nitric acid which is being reduced to NO gas:



whereas the surrounding liquid and accompanying Zn electrode act as the anode compartment, leading to oxidation of the Zn and to a net cell potential of 1.72 V

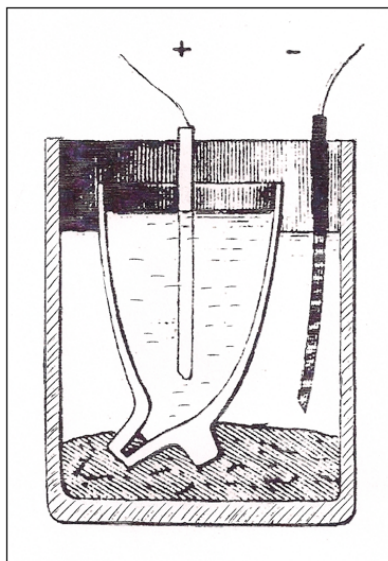


Figure 5. Use of the bowl of a clay pipe as a separator for a small Grove cell.

under standard conditions:



A final chemical use of clay pipes, dating from 1895, is shown in figures 6 and 7 (6). Here they are being used as diffusion membranes for the separation of gases. Since, according to Graham's law of diffusion, the rate of gaseous diffusion is inversely related to the square root of the molecular weight of the gas, dihydrogen gas will diffuse much more rapidly than either the dioxygen or dinitrogen gas in the atmos-

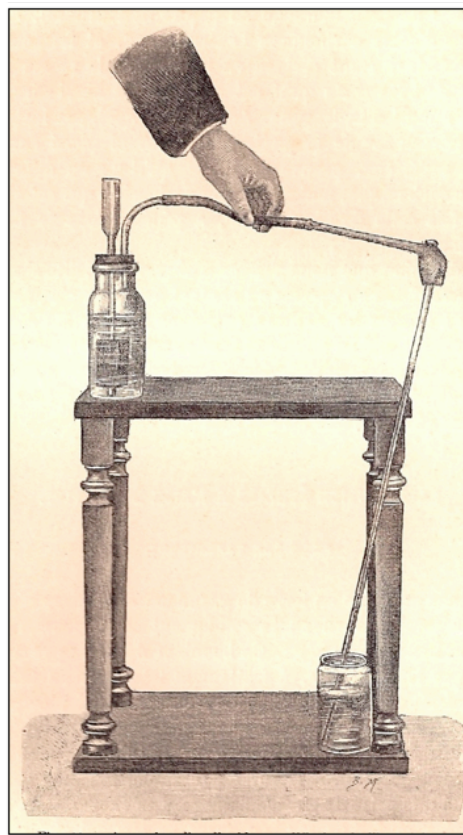


Figure 6. Ascension of a liquid via diffusion of dihydrogen gas.

phere. In figure 6 diffusion of dihydrogen gas out of the clay pipe is used to create a lower pressure above the liquid in the glass tube than that above the liquid in the lower bottle thus causing the liquid to rise in the tube, whereas in figure 7 the situation has been reversed and diffusion of dihydrogen from the surrounding beaker into the clay pipe causes the pressure above the liquid in the bottle to be greater than that outside the bottle, thus driving the liquid out of the bottle in the form of a "marvelous fountain." In both demonstrations it is necessary to seal the mouth of the pipe bowl

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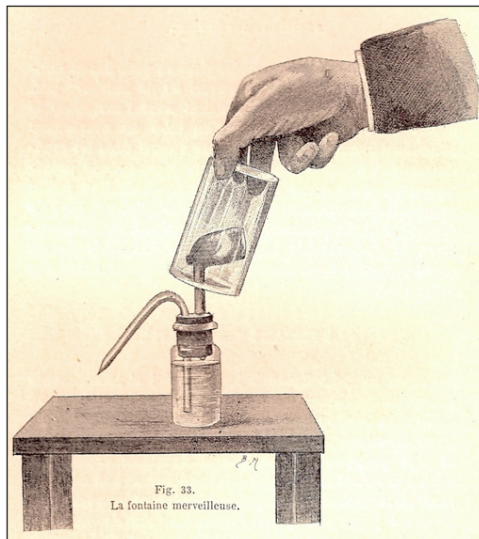


Figure 7. "The marvelous fountain"

using a cork that has been coated with wax.

Günther Beer has found a similar collection of clay tobacco pipes among the holdings of the *Museum der Göttinger Chemie* in Göttingen Germany and has traced references to their uses in the chemical laboratory dating back as far as 1780, including not only those noted above, but far more mundane applications, such as the use of broken pipe stems as stirring rods and sample pipettes (1). However, the earliest reference to clay tobacco pipes in connection with chemical laboratories that I am aware of occurs in the 1680 edition of Johann Becher's *Tripus hermeticus*. This contains

a large foldout plate of essential instruments for the operation of a chemical laboratory of which item 64 shows two clay tobacco pipes, a pouch of tobacco, a candle, an hour glass, and a vial of a poison antidote known as the-riak (7). In this case the clay pipes and tobacco were presumably for the chemist to smoke and not for use as actual apparatus. Interestingly Becher's list of laboratory essentials also included an apron and a lucky rabbit's foot!

References and Notes

1. This column was inspired by a correspondence in 2008 on this subject with Dr. Günther Beer, who is curator of the *Museum der Göttinger Chemie* in Göttingen Germany. See G. Beer, "Die Verwendung von Tonpfeifenteilen als Ersatz-Geräte im chemischen Laboratorium: Eine Literatursammlung von 1787 bis 1947," *Museumbrief*, **2008**, 25, 37-43.
2. K. Thöne, *Chemischen Laboratorium. Anleitung zur Einrichtung und zum Selbstaufbau der erforderlichen Apparate*, Otto Meyer Verlag: Ravensburg, 1947, p. 68.
3. C. R. Gibson, *Chemical Amusements and Experiments*, Seeley, Service & Co: London, 1928, p. 17.
4. J. Partington, *A Text-Book of Inorganic Chemistry*, Macmillan: London, NY, 1925, p. 565.
5. W. R. Grove, "On a Small Voltaic Battery of Extraordinary Energy," *Rep. Brit. Assoc. Adv. Sci.* **1839**, 36-38; also "On a Small Voltaic Battery of Great Energy," *Phil. Mag.*, **1839**, 15 (Series 3), 287-293.
6. F. Faideau, *La chimie amusante*, Librairie Illustrée: Paris, 1892, pp. 72, 74, 76.
7. This plate is reproduced in F. Ferchl, A. Stüssenguth, *A Pictorial History of Chemistry*, Heinemann: London, 1939, p. 117.